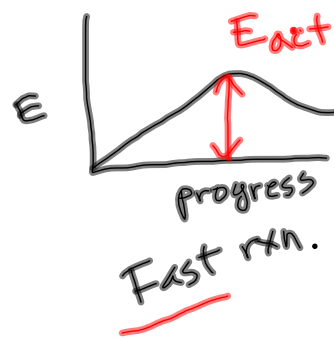
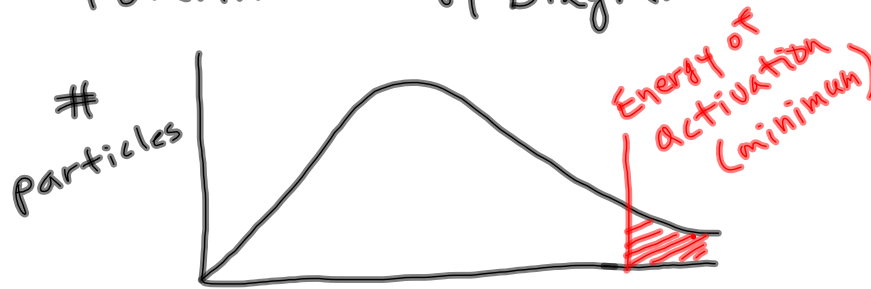


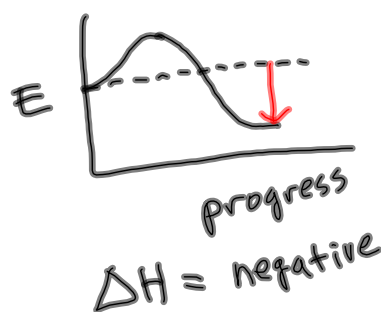
Kinetic Molecular Theory Explains Factors Affecting Rate of Reaction-Continued

3. Nature of Reaction.

- Reacting particles must have a minimum amount of energy to react successfully.
- Potential Energy Diagram



- * Exothermic / Endothermic
- Loss of heat to surroundings e.g. combustion
- Gain of heat from surroundings e.g. plants



Read-Text p.471-477

Kinetic Molecular Theory Explains Factors Affecting Rate of Reaction-Continued

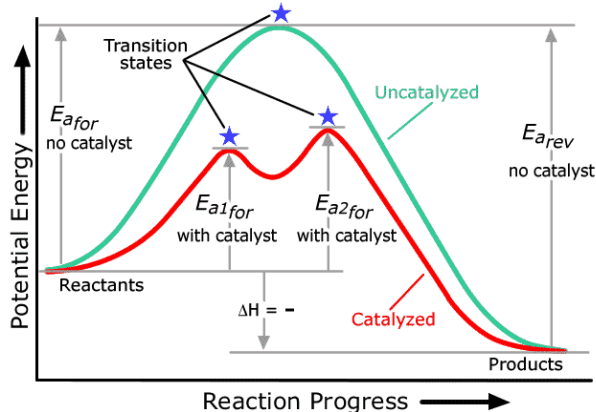
4. Temperature - higher temp. \Rightarrow faster reaction

- higher temp. \rightarrow particles moving faster
- greater chance of collision
- more particles achieve activation E.

\therefore faster reaction Text p.472

5. Catalyst - speeds reaction by lowering E_{act}

- more particles collide successfully (Uncatalyzed reaction continues)
- Faster reaction



- Catalyst is not consumed

Questions - p.476 #1
p.484 #1,2,4