

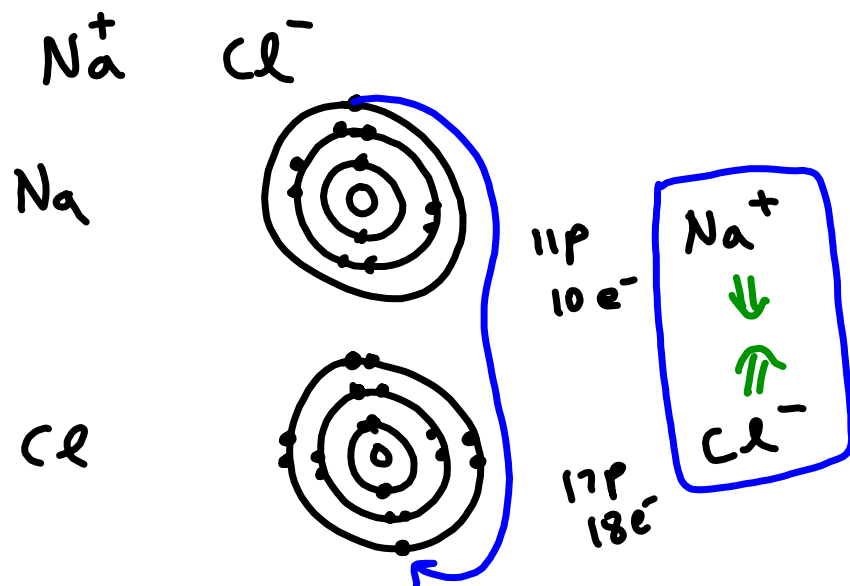
## Chemical Bonding-Ionic

Ions - charged particles

Ionic compounds conduct electricity

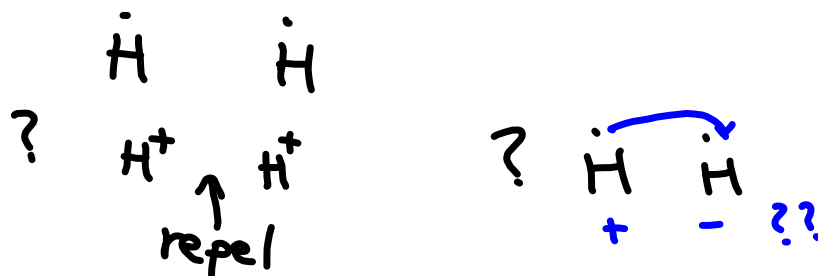
High melting point - strong attraction

e.g. NaCl - sodium chloride



## Covalent Bonding

Consider-  $H_2$  (gas) ? bonding



What if the atoms share electrons?



Each H atom "has"  $2e^-$   
 opposite charges attract  
 simultaneously

Covalent Bonding

between two  
 non-metals

## Chemical Formulas and Names

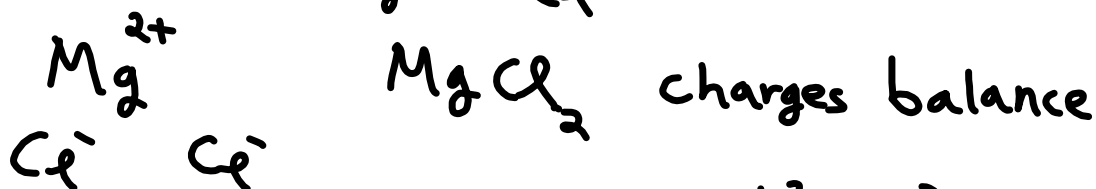
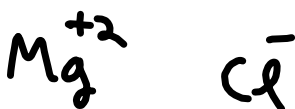
\* Given a name, write a formula

A. Ionic- Sodium chloride

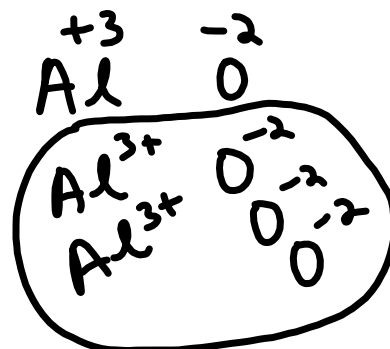
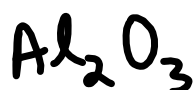


$\text{NaCl}$  - balanced charges

2. Magnesium chloride

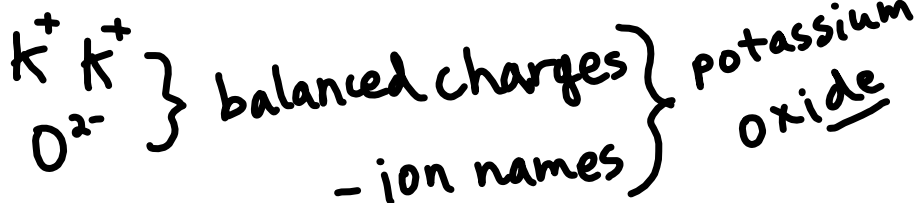


3. Aluminum oxide



Given formula → name

e.g. 1.  $\text{K}_2\text{O}$  → name?



2. Complex ion-  $\text{Ca}(\text{NO}_3)_2$       $\text{Ca}^{2+}$       $\text{NO}_3^-$   
 Calcium nitrate      $\text{NO}_3^-$

Questions- p.26 #13-16, 17, 19  
 (choose)

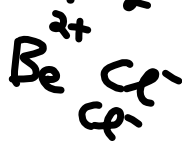
## Naming Molecular Compounds (Covalent)

Examples -

1.  $\text{CO}_2$  carbon dioxide
2.  $\text{CO}$  carbon monoxide
3.  $\text{N}_2\text{O}_5$  dinitrogen pentoxide
4. Write - formula for sulfur trioxide



5.  $\text{BeCl}_2 \rightarrow$  name



Questions p. 24 #9-11

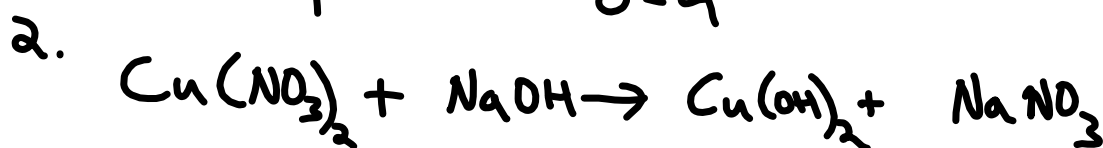
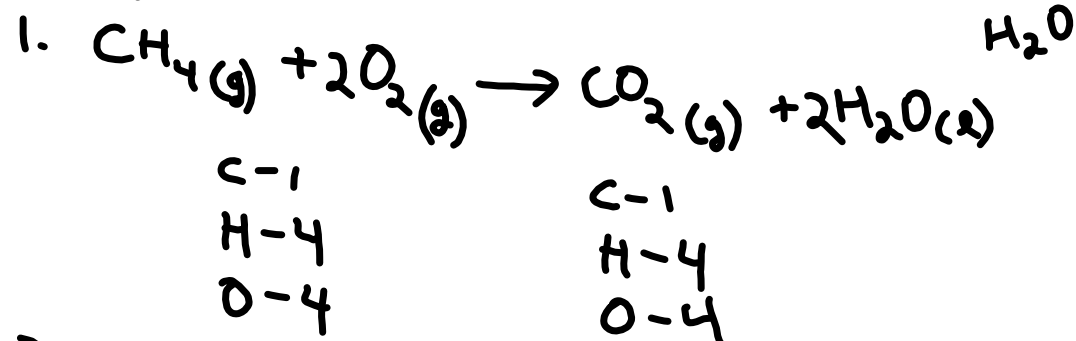
## Chemical Reactions and Equations

Text- p.29-30

Conservation of mass - not created or destroyed

\* Given an unbalanced equation, balance it using coefficients.

examples-



Quest. p.31 #21-23

balance only