Faraday's Laws Example - Time

How much time is needed to deposit 15.8g Ag using a 25.0 A current?

1. $15.8\,g\,Ag \times \frac{1\,mol\,Ag}{107.87\,g} = 0.146\,mol\,Ag$

2. $Ag^+ + e^- \longrightarrow Ag$

$0.146\,mol\,Ag \times \frac{1\,mol\,e^-}{1\,mol\,Ag} = 0.146\,mol\,e^-$

3. $0.146\,mole^- \times \frac{96\,500\,C}{1\,mole^-} = 14\,135\,C$

also 14135 A-S

4. $\frac{14\,135\,A\text{-}S}{25.0\,A} = 565\,sec.$

p. 793 #22 - time
#23 - A