Try some practice problems on your own! Refer to the tutorials if you get stuck.

1. PbSO4 (s) <---> Pb2+ (aq) + SO42- (aq)

[Pb2+] = 2.0 x 10-5 M   Find the value of Ksp.

2. Solid barium phosphate dissolves into its respective ions at 25oC. If you are given that

[Ba2+] = 1.5 x 10-2 M, find the value of Ksp.

3. In a solution of 0.200M NaI, solid AgI was placed in. Calculate the solubility of AgI (Ksp = 1.5 x 10-16)

4. Solid barium phosphate dissolves into its respective ions at 25oC. Given that [Ba2+] = 1.5 x 10-2 M, find the value of Ksp.

5. Solid silver chloride dissolves into its respective ions at 25oC. Given that [Ag+] = 3.5 x 10-2 M, find the value of Ksp.

6. Solid aluminum hydroxide dissolves into its respective ions at 25oC. Given that  [OH- ] = 1.5 x 10-5 M, find the value of Ksp.

7. In a solution of 0.200M NaI, solid AgI was placed in. Calculate the solubility of AgI (Ksp = 1.5 x 10-16)

8. In a solution of 0.100M Pb(NO3)2, solid PbCO3 was placed in. Calculate the solubility of PbCO3

(Ksp = 3.3×10-14)

9. In a solution of 0.500M Na2S, solid FeS was placed in. Calculate the solubility of FeS

(Ksp = 3.7×10-19)

10. Calculate the solubility of AlPO4. (Ksp = 9.8 x 10-21)

11. Calculate the solubility of CdCO3. (Ksp = 1.0 x 10-12)

12. Calculate the solubility of MgF2. (Ksp = 6.4 x 10-9)