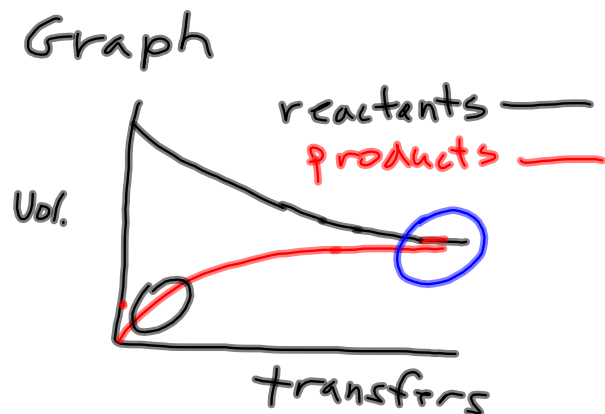


Equilibrium Analogy Notes



1. reaction slows over time (react \rightarrow prod.)
3. reaction is faster at beginning
4. equilibrium is found near the "end" - graph is flat
5. At equilibrium, rates are equal.
This means the rate of collisions of reactant particles is the same as the rate of collisions of product particles.
6. There are always some reactants - not all convert to products. This happens because of the "reverse" reaction.