

Temperature and Rate of Reaction

Higher Temperature = faster rate

Higher temp → particles move faster

→ higher chance of colliding,

→ more particles achieve activation E

∴ Reaction is faster

Catalyst and Rate of Reaction

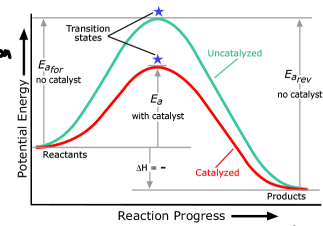
Catalyst speeds reaction by lowering E_{act} and is not used up.

Note- the uncatalyzed reaction still goes on.

Questions- text

p. 476 #1

p. 484 #1,2,4

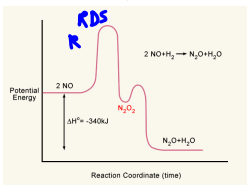
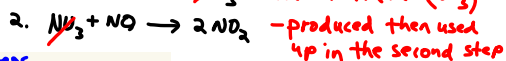
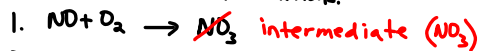


Reaction Mechanisms and Rate-Determining Step

Text p. 477

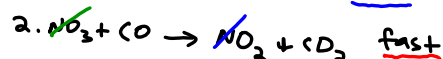
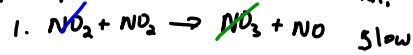


This reaction is made up of several steps called elementary reactions.



This could represent a two-step reaction
Step 1 is rate-determining

Example - Write the overall rxn.



Step 1 (slow) is rate determining

Text p. 478 #5-8

p. 486 #10,11

Read p. 479 →