Test 1 Outline

1. Models of the atom (by name)
2. Bohr model + quantum mechanical
   - Bohr model → orbits
   - quantum mechanical → orbitals
3. Explain how spectral lines are evidence.
4. Write & explain e^- configuration
   (representative elements "A")
   - d-block
5. Diagonal rule – periodic table
6. Valence electrons
7. Trends in the periodic table.
   (atomic radius & ionization E

Factors Affecting Rate of Reaction

Read: Text 467

1. Temperature - Higher temp. = faster
   e.g. storing food in the refrigerator to slow spoiling.

2. Concentration of reactants
   Higher Concentration ⇒ faster reaction

3. Surface area -
   more surface area ⇒ faster reaction
   e.g. firewood - split to smaller pieces
   ↑ greater S.A.

4. Nature of reaction -
   some reactions are fast
   " " " slow

5. Catalyst - substances which speed up a rxn.
   e.g. enzymes
   car exhaust system - cat. converter