

Test 1 Outline

1. Models of the atom. (by name)
2. Bohr + Quantum Mechanical
3. Spectral lines as evidence.
4. e<sup>-</sup> configuration- (representative elements)  
Write + explain
5. Diagonal rule
6. Valence electrons + dot diagrams
7. Trends in the periodic table
  - atomic radius
  - ionization E
  - predict or explain

Rate of Reaction- Definition

Examples- speed signs -  $100 \frac{\text{km}}{\text{h}} \leftarrow \frac{\Delta \text{ distance}}{\Delta \text{ time}}$

- lab demonstration

$\frac{\Delta \text{ Size of balloon}}{\Delta \text{ time}}$

Reactions-  $\frac{\Delta \text{ mass of product/reactant}}{\Delta \text{ time}}$

Also- colour,  $\Delta \text{ time}$ , temperature, voltage, volume

Read text- p.462, 463

Read lab handout.

Factors Affecting Rate of Reaction