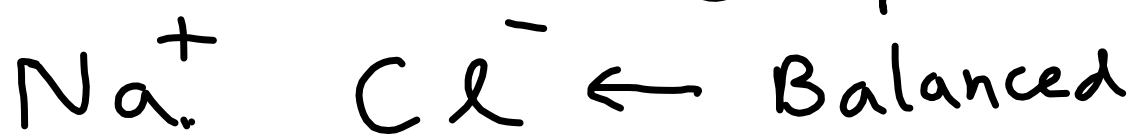


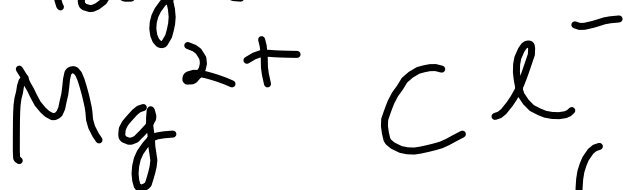
Naming/Writing Formulas-Ionic

Given a name \rightarrow formula

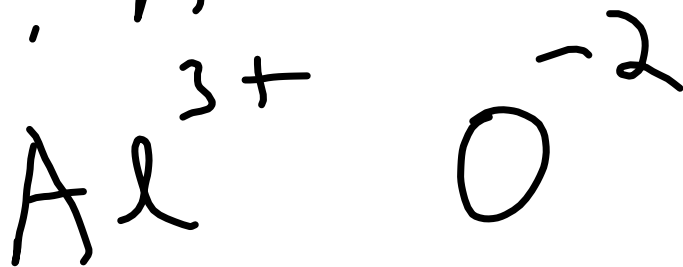
e.g. 1. Sodium chloride \leftarrow
 compound



2. Magnesium chloride



3. Aluminum Oxide



* Given formula \rightarrow chemical name

e.g. 1. K_2O K^+ O^{2-} balance
 K^+ O^{2-} K^+

potassium oxide $2+$ $(1-)_2$

2. Complex ion $Ca(NO_3)_2$

Calcium nitrate

P. 26 Questions - ~~#~~ 13-16

17, 19

(Choose)

Covalent Bonding

Consider - H_2 (gas)

$\dot{H} \quad \dot{H} \quad ?$ bonding

Table of
ions

$H^+ \quad H^+$

Proposed - sharing e^-

Atoms "try" to become more stable
Achieve a noble gas arrangement

$H \ominus H$ Opposite charges attract
simultaneously - bond

$+ \quad - \quad +$

Covalent Bond - between two or
more non-metal atoms

Naming Molecular Compounds (Covalent)

Examples -

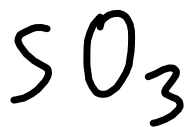
1. CO_2 - Carbon dioxide
prefix

2. CO - carbon monoxide

text p. 23 - table (prefixes)
no balance "+" "-"

3. N_2O_5
↳ dinitrogen pentoxide

4. Write the formula for
sulfur trioxide



Questions
p. 24
#9-11

Chemical Reactions and Equations

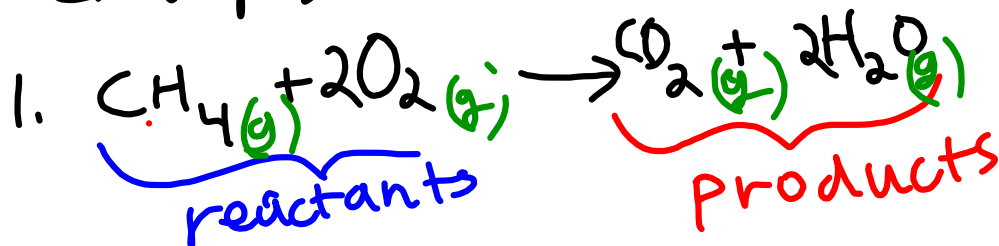
Text - p. 29
30

* Matter is conserved.

↓
not created
or destroyed

* Given an unbalanced equation,
balance using coefficients

Examples -

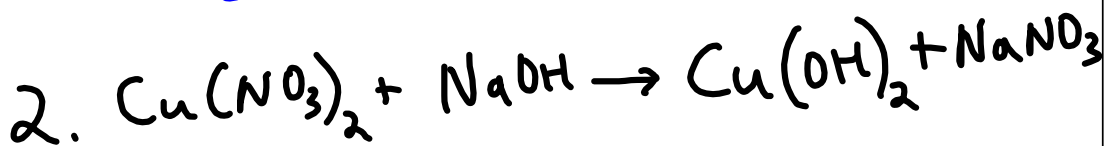


Reactants

C - 1
H - 4
O - 4

Products

C - 1
H - 4
O - 4



Bal.
only

Quest. - p. 31
21-23